

1. Plot a Curve of pH against ml of HCl added.
2. determine the practical buffer capacity in the acid directions, from the graph and the formula.

Acetate buffer	practical capacity (from the formula)	practical capacity (from the curve)
0.1M Acetate buffer		
0.2M Acetate buffer		

In the Discussion

- Which buffer has more capacity, and why?
- Compare the value of the buffer capacity you got from the curve and formula.
- Discuss and compare the titration curves of the two buffers
- Do you think your buffer has a larger capacity for acid or base? Why?