

# **Chapter 1**

## **INTRODUCTION AND BASIC CONCEPTS**

# Objectives

- Identify the unique vocabulary of thermodynamics
- Review the metric SI and the English unit systems.
- Explain the basic concepts of thermodynamics
  - **Examples:** system, state, state postulate, equilibrium, process, and cycle.
- Review concepts of temperature, temperature scales, pressure, and absolute and gage pressure.
- Introduce a simple systematic problem-solving technique.

# THERMODYNAMICS AND ENERGY

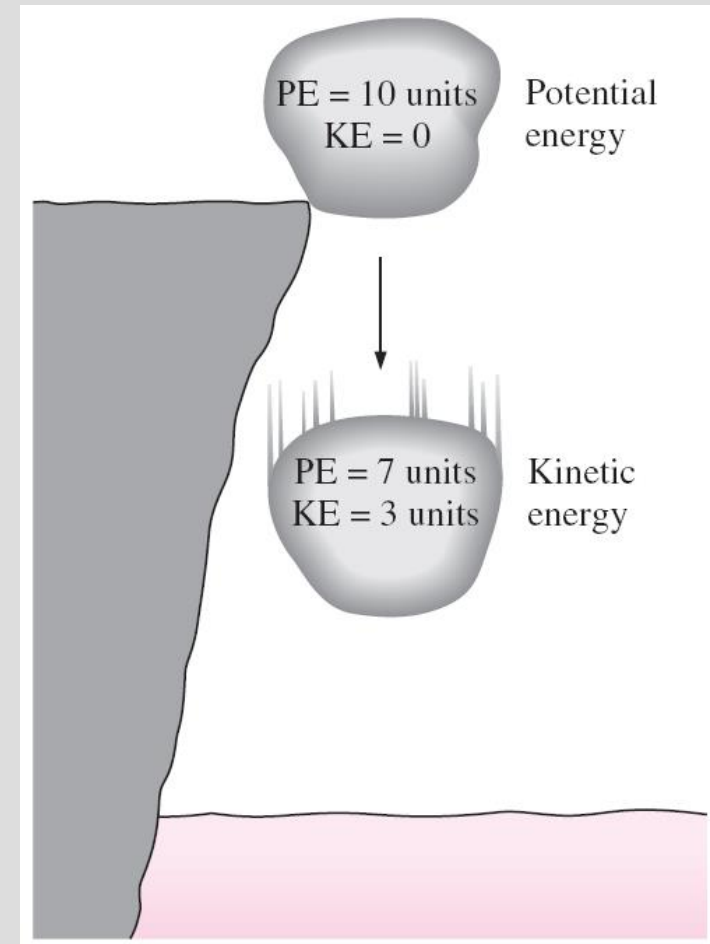
- **Thermodynamics:** The science of *energy*.
- **Energy:** The ability to cause changes – The capacity to do work.
- The name *thermodynamics* stems from the Greek words *therme* (heat) and *dynamis* (power).

# CONSERVATION OF ENERGY

- **Conservation of energy principle:** Energy can change from one form to another but the total amount of energy remains constant.
- Energy cannot be created or destroyed.



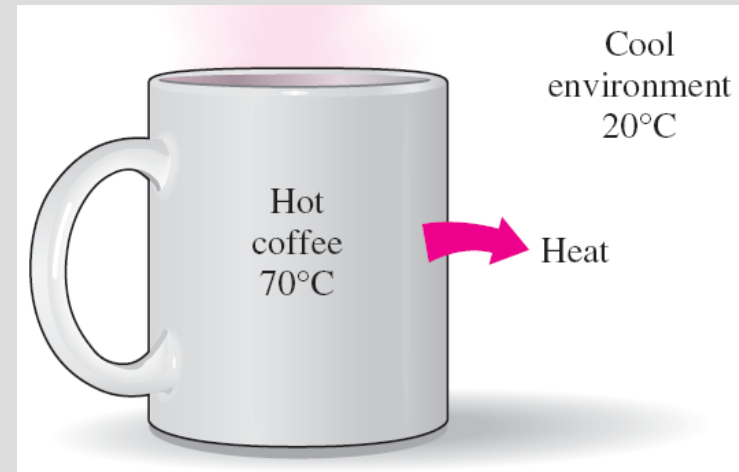
Conservation of energy principle for the human body.



Energy cannot be created or destroyed; it can only change forms (the first law).

# LAWS OF THERMODYNAMICS

- **The first law of thermodynamics:** An expression of the conservation of energy principle.
- The first law asserts that *energy* is a thermodynamic property.
- **The second law of thermodynamics:** It asserts that energy has *quality* as well as *quantity*, and actual processes occur in the direction of decreasing quality of energy.

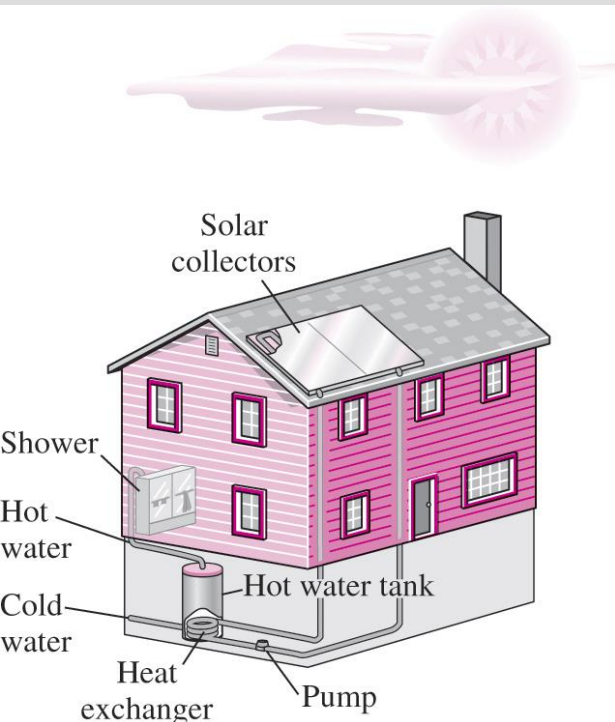


Heat flows in the direction of decreasing temperature.

# TYPES OF THERMODYNAMIC ANALYSIS

- **Classical thermodynamics:** A macroscopic approach to the study of thermodynamics that does not require a knowledge of the behavior of individual particles.
- It provides a direct and easy way to the solution of engineering problems
- **Statistical thermodynamics:** A microscopic approach, based on the average behavior of large groups of individual particles.
- Only classical thermodynamics will be studied in this course.

# Application Areas of Thermodynamics





# Application Areas of Thermodynamics





# IMPORTANCE OF DIMENSIONS AND UNITS

- Any physical quantity can be characterized by **dimensions**.
- The magnitudes assigned to the dimensions are called **units**.
- Some basic dimensions such as mass  $m$ , length  $L$ , time  $t$ , and temperature  $T$  are selected as **primary** or **fundamental dimensions**,
- Other dimensions such as velocity  $V$ , energy  $E$ , and volume  $V$  are expressed in terms of the primary dimensions and are called **secondary dimensions**, or **derived dimensions**.

**TABLE 1–1**

The seven fundamental (or primary) dimensions and their units in SI

Dimension	Unit
Length	meter (m)
Mass	kilogram (kg)
Time	second (s)
Temperature	kelvin (K)
Electric current	ampere (A)
Amount of light	candela (cd)
Amount of matter	mole (mol)

# METRIC vs. ENGLISH SYSTEM OF UNITS

- **Metric SI system:** A simple and logical system based on a decimal relationship between the various units.
- **English system:** It has no apparent systematic numerical base, and various units in this system are related to each other rather arbitrarily.

**TABLE 1–2**

Standard prefixes in SI units

Multiple	Prefix
$10^{12}$	tera, T
$10^9$	giga, G
$10^6$	mega, M
$10^3$	kilo, k
$10^2$	hecto, h
$10^1$	deka, da
$10^{-1}$	deci, d
$10^{-2}$	centi, c
$10^{-3}$	milli, m
$10^{-6}$	micro, $\mu$
$10^{-9}$	nano, n
$10^{-12}$	pico, p

# Some SI and English Units

$$1 \text{ lbm} = 0.45359 \text{ kg}$$

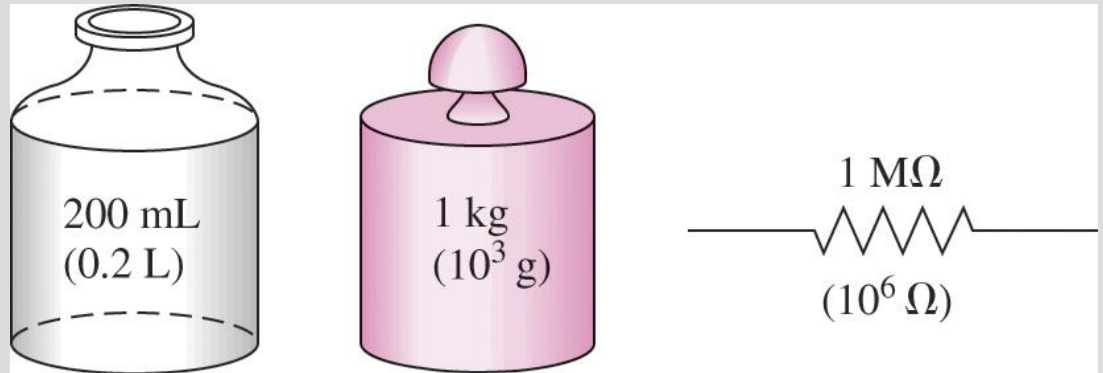
$$1 \text{ ft} = 0.3048 \text{ m}$$

$$\text{Work} = \text{Force} \times \text{Distance}$$

$$1 \text{ J} = 1 \text{ N} \cdot \text{m}$$

$$1 \text{ cal} = 4.1868 \text{ J}$$

$$1 \text{ Btu} = 1.0551 \text{ kJ}$$



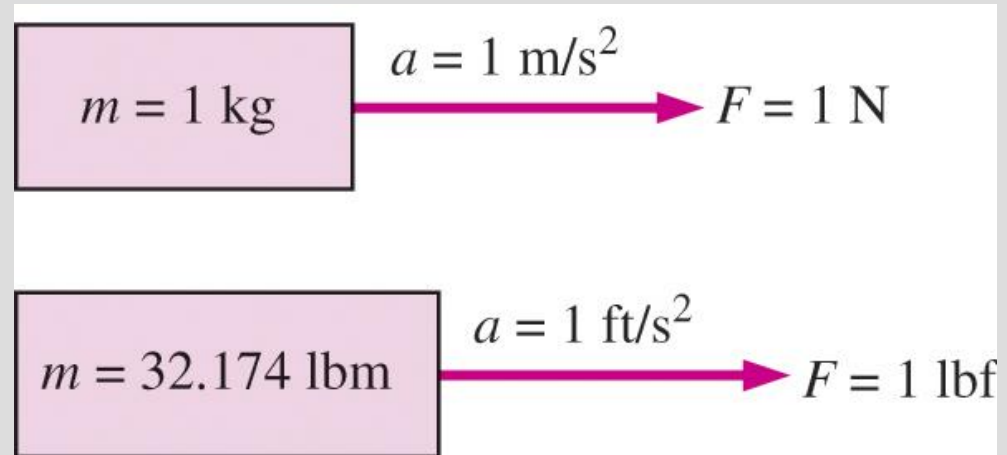
The SI unit prefixes are used in all branches of engineering.

$$\text{Force} = (\text{Mass})(\text{Acceleration})$$

$$F = ma$$

$$1 \text{ N} = 1 \text{ kg} \cdot \text{m/s}^2$$

$$1 \text{ lbf} = 32.174 \text{ lbm} \cdot \text{ft/s}^2$$



The definition of the force units.

# Dimensional homogeneity

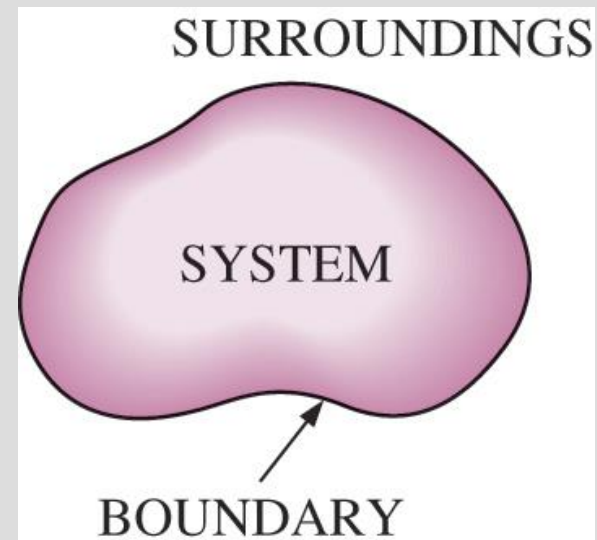
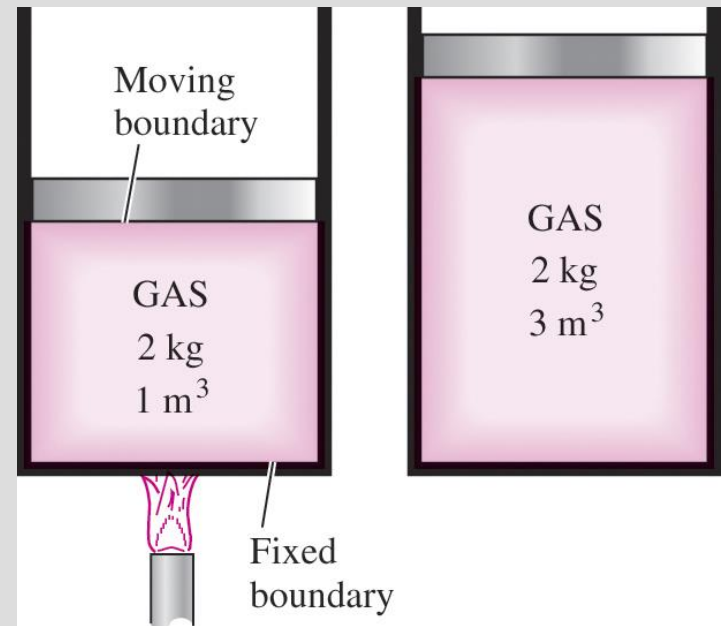
- All equations must be dimensionally **homogeneous**.
- To be dimensionally homogeneous, all the terms in an equation must have the same unit.

## EXAMPLE

A tank is filled with oil whose density is 0.85 kg/L. The volume of the tank is 2 m<sup>3</sup>. What is the mass of oil in the tank?

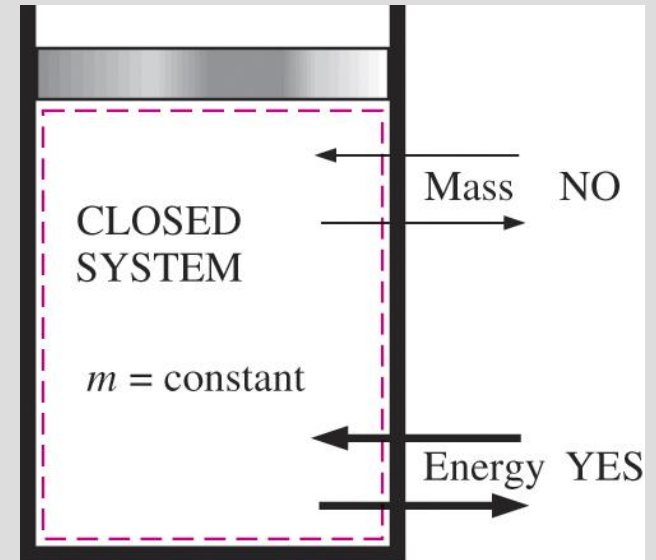
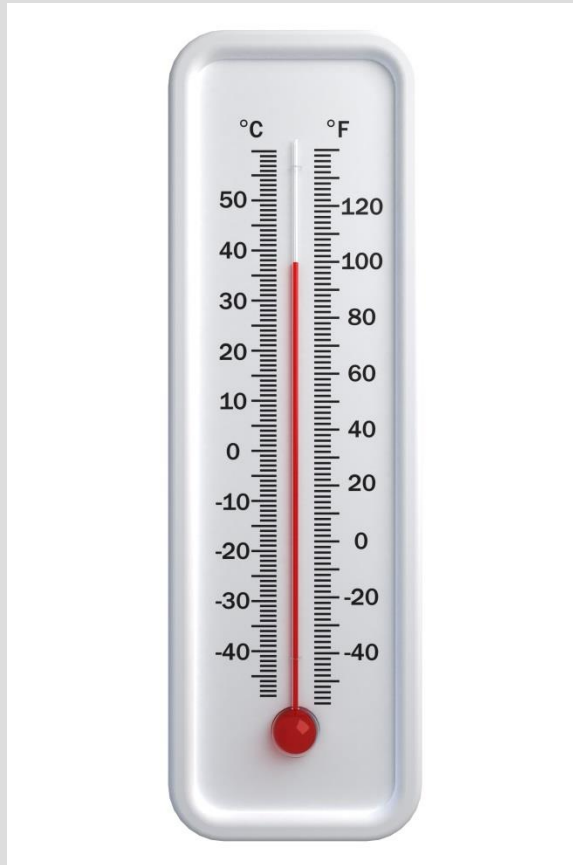
# SYSTEMS

- **System:** A quantity of matter or a region in space chosen for study.
- **Surroundings:** The mass or region outside the system
- **Boundary:** The real or imaginary surface that separates the system from its surroundings.
- The boundary of a system can be *fixed* or *movable*.
- Systems may be considered to be *closed* or *open*.



# CLOSED SYSTEMS

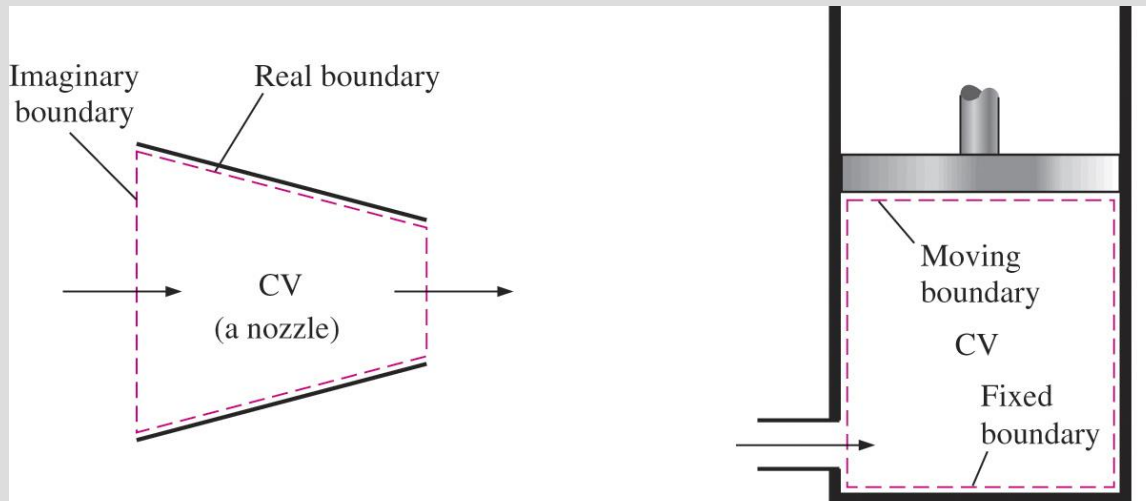
- **Closed system (Control mass):**  
A fixed amount of mass, and no mass can cross its boundary.





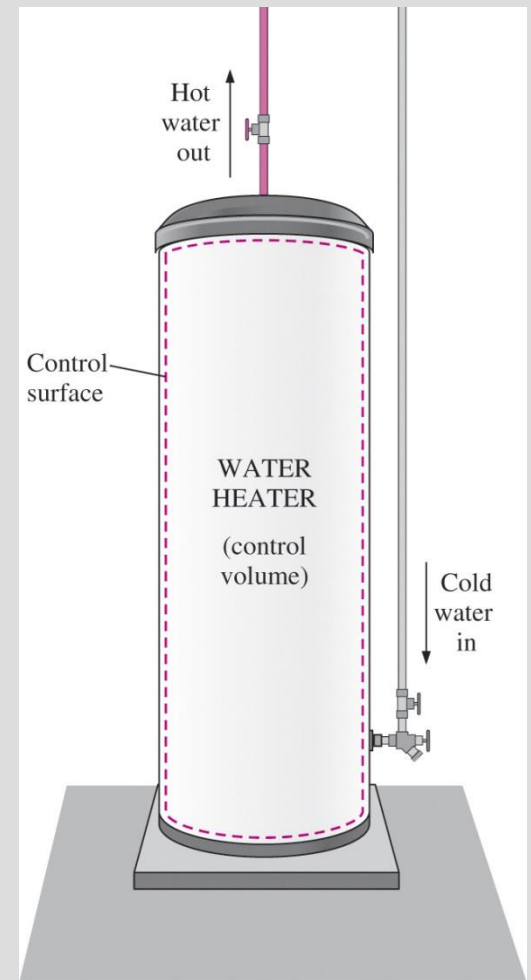
# OPEN SYSTEMS

- **Open system (control volume):** A properly selected region in space.
- It usually encloses a device that involves mass flow such as a compressor, turbine, or nozzle.
- Both mass and energy can cross the boundary of a control volume.
- **Control surface:** The boundaries of a control volume. It can be real or imaginary.



(a) A control volume with real and imaginary boundaries

(b) A control volume with fixed and moving boundaries



An open system (a control volume) with one inlet and one exit.

# PROPERTIES OF A SYSTEM

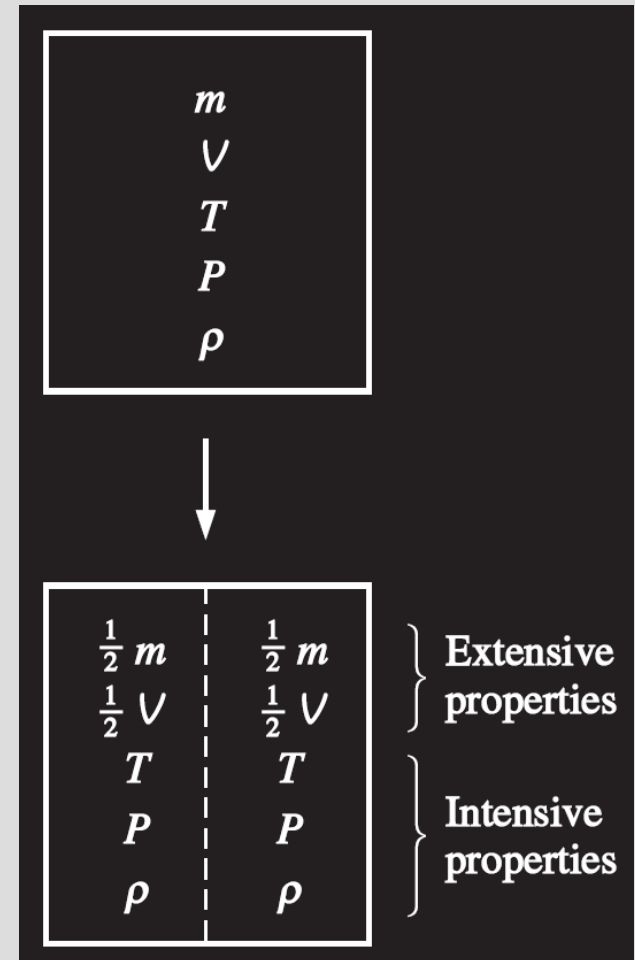
- **Property:** Any characteristic of a system.
- Some familiar properties are pressure  $P$ , temperature  $T$ , volume  $V$ , and mass  $m$ .

**OTHER EXAMPLES?**

# INTENSIVE AND EXTENSIVE PROPERTIES

- Properties are considered to be either *intensive* or *extensive*.
- **Intensive properties:** Those that are independent of the mass of a system, such as temperature, pressure, and density.
- **Extensive properties:** Those whose values depend on the size—or extent—of the system.

## EXAMPLES?



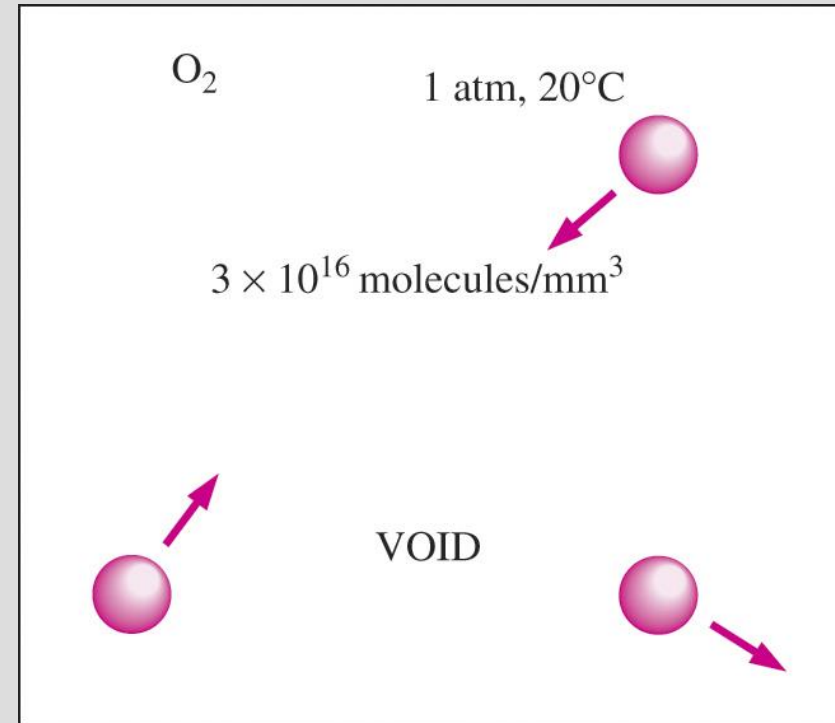
Criterion to differentiate intensive and extensive properties.

# SPECIFIC PROPERTIES

- **Specific Property:** Extensive property per unit mass
- Specific properties are denoted by lower case letters
- **Examples:** specific volume ( $v$ ), specific energy ( $e$ )
- Specific properties are intensive properties

# Continuum

- Matter is made up of atoms that are widely spaced in the gas phase.
- Yet a substance can be treated as a continuum (homogeneous with no holes) because of the very large number of molecules even in a small volume.
- This assumption is valid as long as the size of the system we deal with is large relative to the space between the molecules.
- This is the case in practically all problems.
- In this text we will limit our consideration to substances that can be modeled as a continuum.



# DENSITY AND SPECIFIC GRAVITY

## Density

$$\rho = \frac{m}{V} \quad (\text{kg/m}^3)$$

## Specific volume

$$\nu = \frac{V}{m} = \frac{1}{\rho}$$

**Specific gravity:** The ratio of the density of a substance to the density of some standard substance at a specified temperature (usually water at 4°C).

$$SG = \frac{\rho}{\rho_{\text{H}_2\text{O}}}$$

TABLE 1–3

Specific gravities of some substances at 0°C

Substance	SG
Water	1.0
Blood	1.05
Seawater	1.025
Gasoline	0.7
Ethyl alcohol	0.79
Mercury	13.6
Wood	0.3–0.9
Gold	19.2
Bones	1.7–2.0
Ice	0.92
Air (at 1 atm)	0.0013

$$V = 12 \text{ m}^3$$
$$m = 3 \text{ kg}$$



$$\rho = 0.25 \text{ kg/m}^3$$

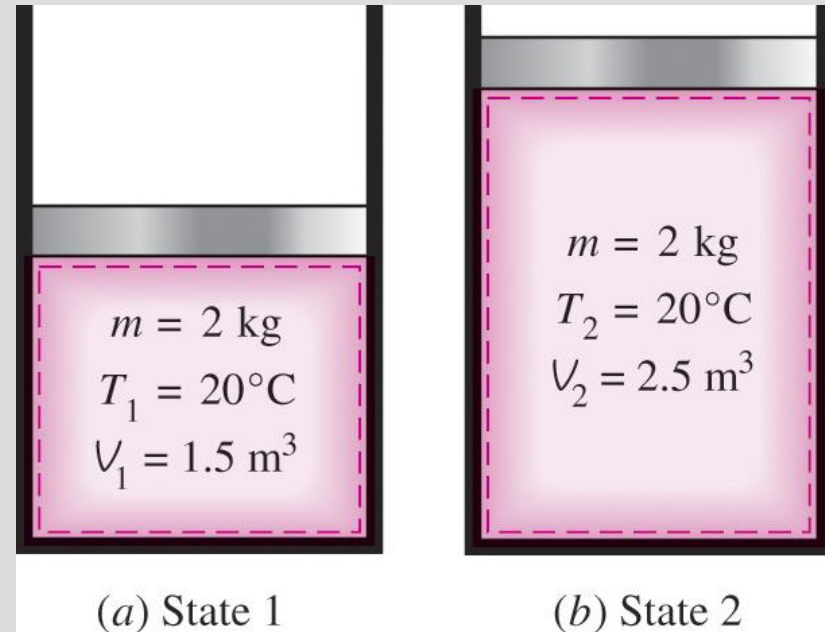
$$\nu = \frac{1}{\rho} = 4 \text{ m}^3/\text{kg}$$

Density is mass per unit volume;  
specific volume is volume per unit mass.



# THERMODYNAMICS STATE

- Consider a system **not** undergoing any change.
- At this point, all the properties can be measured or calculated throughout the entire system.
- This gives us a set of properties that completely describes the condition of the system.
- This set of properties describe the thermodynamic state of the system.
- At a given state, all the properties of a system have fixed values.
- If the value of even one property changes, the state will change to a different one.



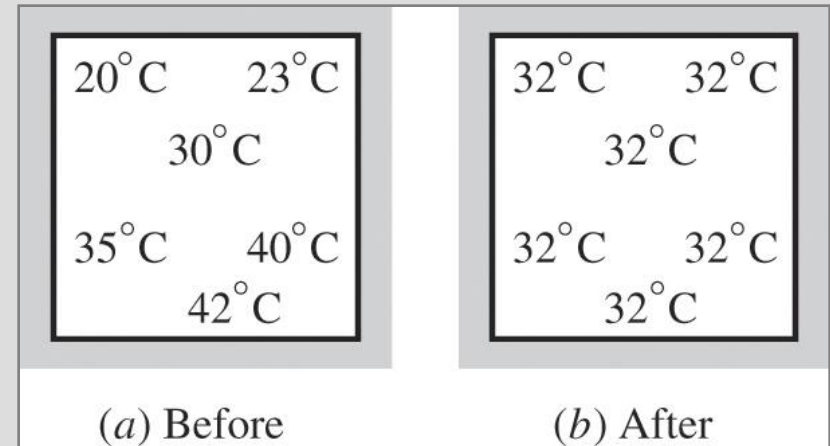
A system at two different states.

# THERMODYNAMIC EQUILIBRIUM

- **Equilibrium:** A condition of balance.
- In equilibrium, there are no unbalances that cause changes in the system.
- A system is in thermodynamic equilibrium if it has:
  - ✓ Thermal equilibrium
  - ✓ Mechanical equilibrium
  - ✓ Phase equilibrium
  - ✓ Chemical equilibrium

# ELEMENTS OF THERMODYNAMIC EQUILIBRIUM

- **Thermal equilibrium:** If the temperature is the same throughout the entire system.
- **Mechanical equilibrium:** If there is no change in pressure at any point of the system with time.
- **Phase equilibrium:** If a system involves two phases and when the mass of each phase reaches an equilibrium level and stays there.
- **Chemical equilibrium:** If the chemical composition of a system does not change with time, that is, no chemical reactions occur.



A closed system reaching thermal equilibrium.

# The State Postulate

- The number of properties required to fix the state of a system is given by the **state postulate**:
  - ✓ *The state of a simple compressible system is completely specified by two independent, intensive properties.*
- **Simple compressible system**: If a system involves no electrical, magnetic, gravitational, motion, and surface tension effects.



The state of nitrogen is fixed by two independent, intensive properties.

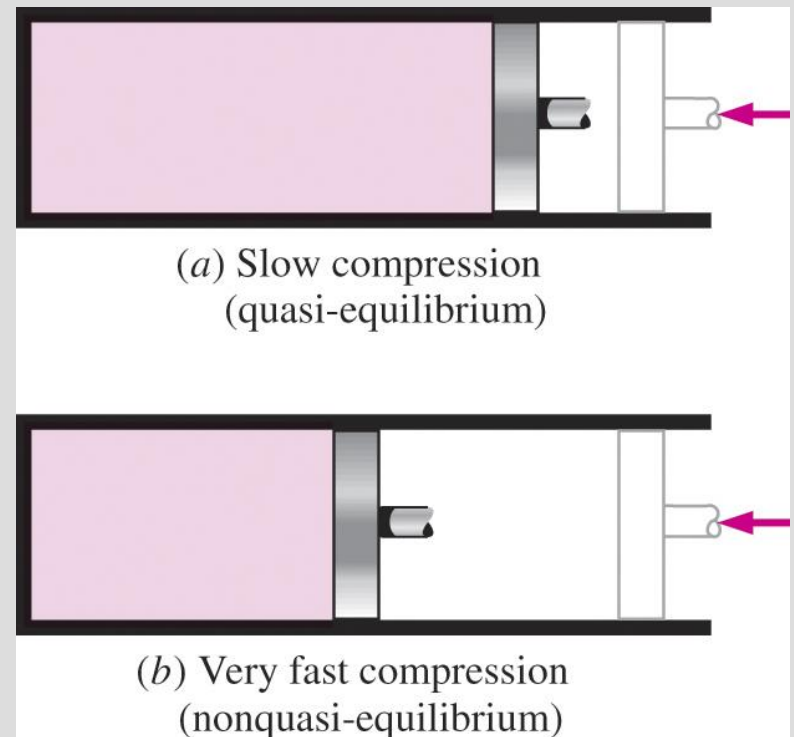
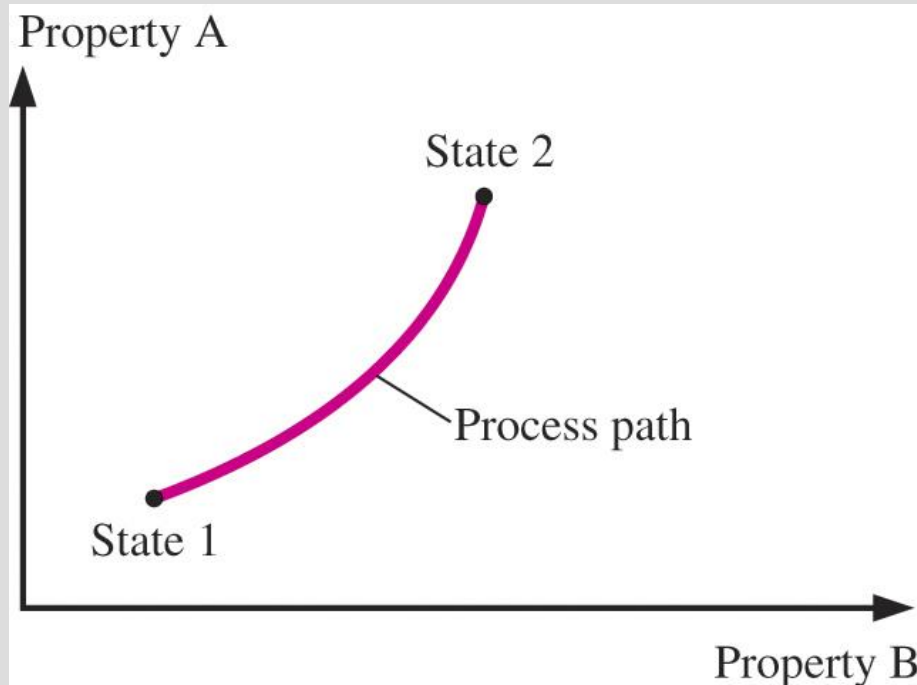
# PROCESSES AND CYCLES

**Process:** Any change that takes a system from one equilibrium state to another.

**Path:** The series of states through which a system passes during a process.

To describe a process completely, one should specify the initial and final states, as well as the path it follows, and the interactions with the surroundings.

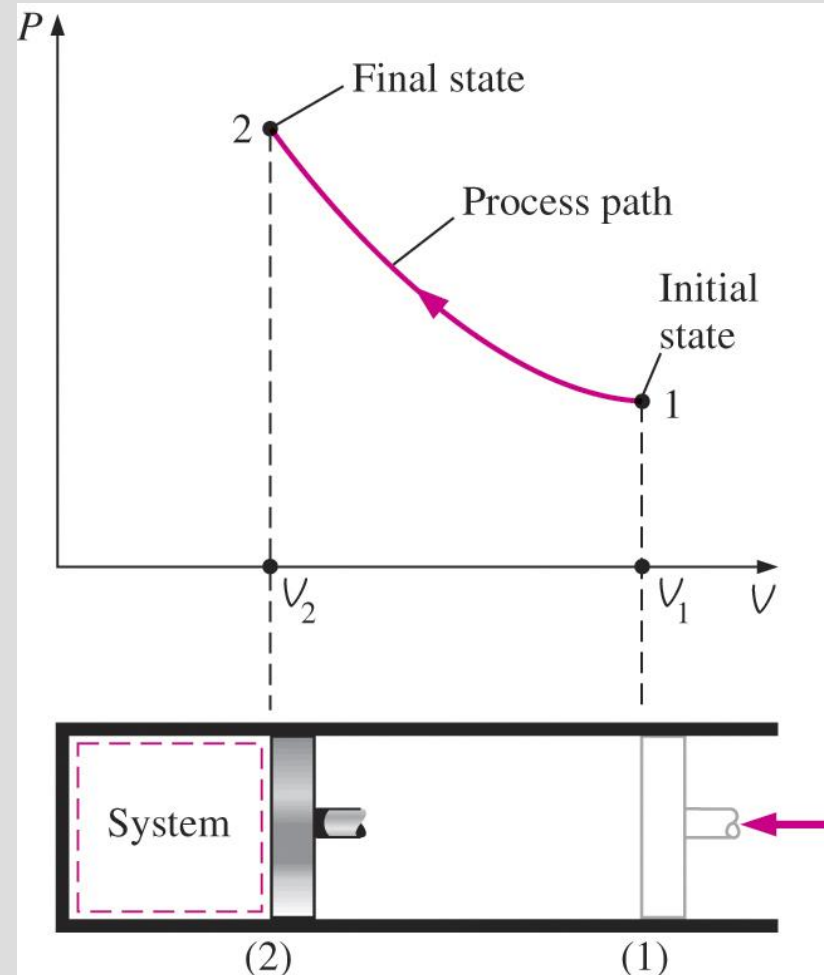
**Quasistatic or quasi-equilibrium process:** When a process proceeds in such a manner that the system remains infinitesimally close to an equilibrium state at all times.



- Process diagrams are very useful in visualizing the processes.
- The prefix *iso-* is often used to designate a process for which a particular property remains constant.

## EXAMPLES

- **Isothermal process:** A process during which the temperature  $T$  remains constant.
- **Isobaric process:** A process during which the pressure  $P$  remains constant.
- **Isochoric (or isometric) process:** A process during which the specific volume  $v$  remains constant.
- **Cycle:** A process during which the initial and final states are identical.

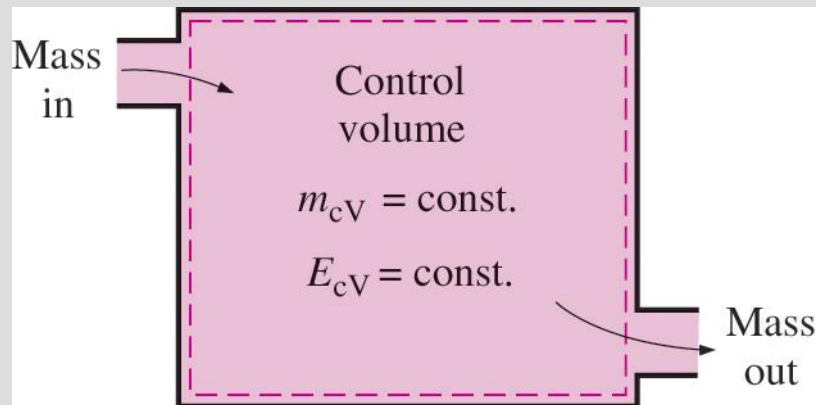


The  $P$ - $V$  diagram of a compression process.

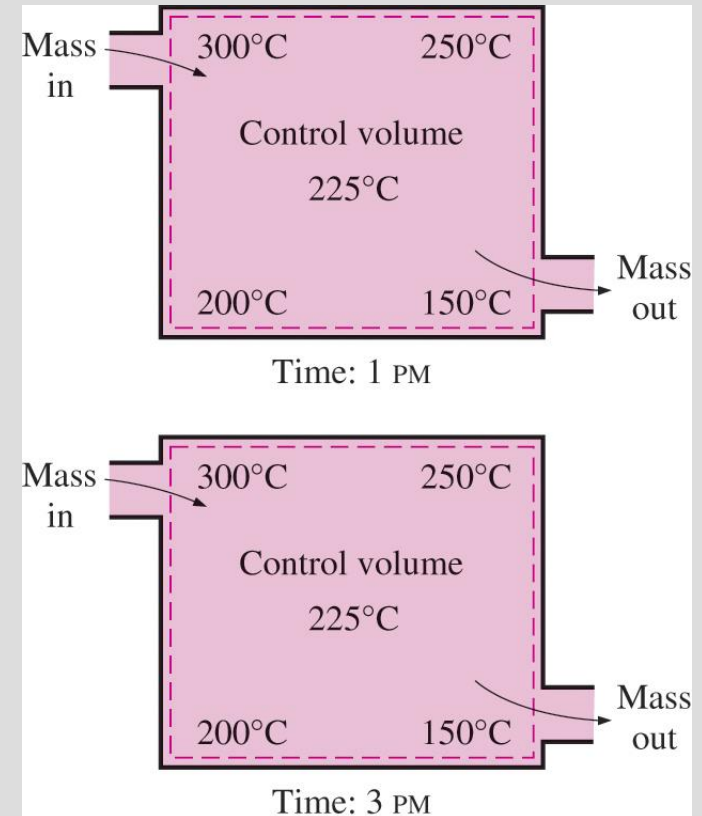


# The Steady-Flow Process

- The term *steady* implies *no change with time*. The opposite of steady is *unsteady*, or *transient*.
- Steady-flow process:** A process during which a fluid flows through a control volume steadily.



Under steady-flow conditions, the mass and energy contents of a control volume remain constant.



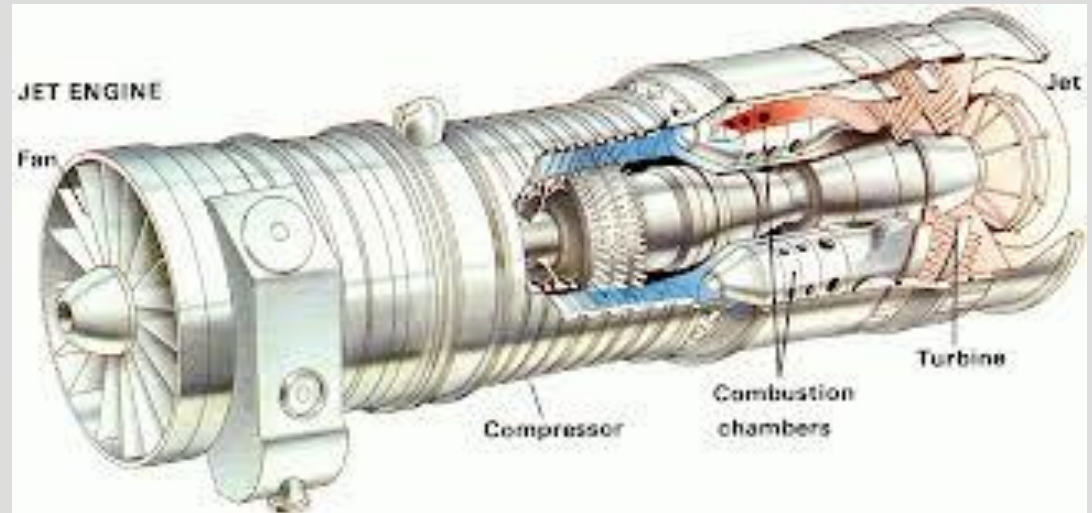
During a steady-flow process, fluid properties within the control volume may change with position but not with time.

# Steady-Flow Devices

- A large number of engineering devices operate for long periods of time under the same conditions, and they are classified as *steady-flow devices*.
- EXAMPLES: turbines, pumps, boilers, condensers, and heat exchangers or power plants or refrigeration systems.



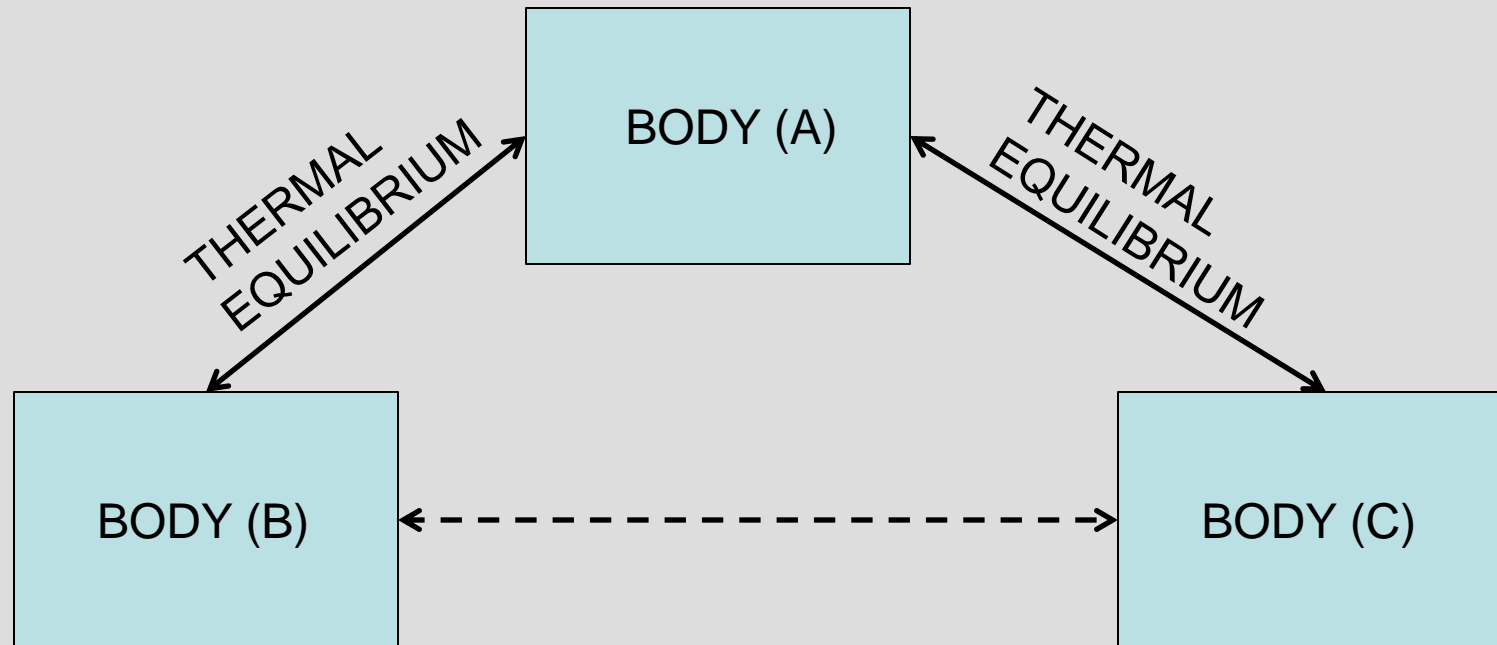
Condenser



Jet Engine

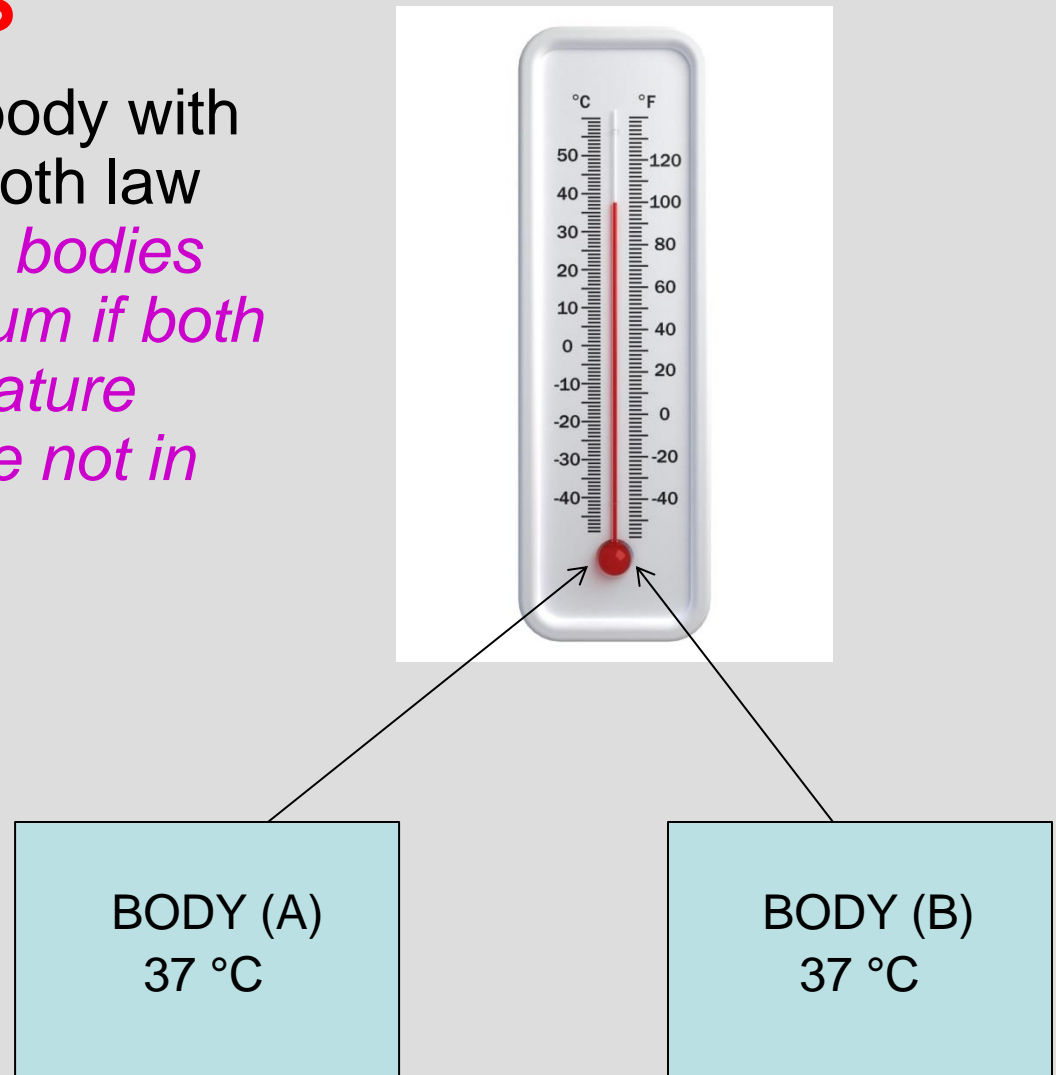
# THE ZEROth LAW OF THERMODYNAMICS

- **The zeroth law of thermodynamics:** If two bodies are in thermal equilibrium with a third body, they are also in thermal equilibrium with each other.



# TEMPERATURE AND THE ZEROth LAW OF THERMODYNAMICS

- By replacing the third body with a thermometer, the zeroth law can be restated as *two bodies are in thermal equilibrium if both have the same temperature reading even if they are not in contact.*

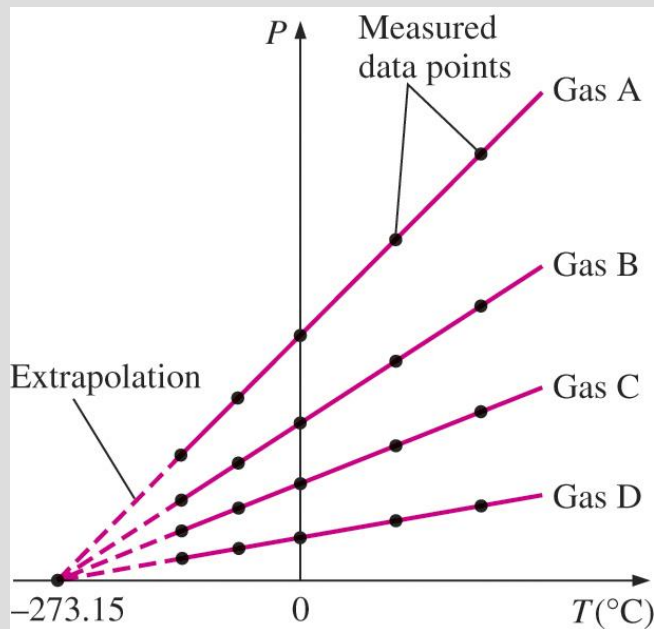


# Temperature Scales

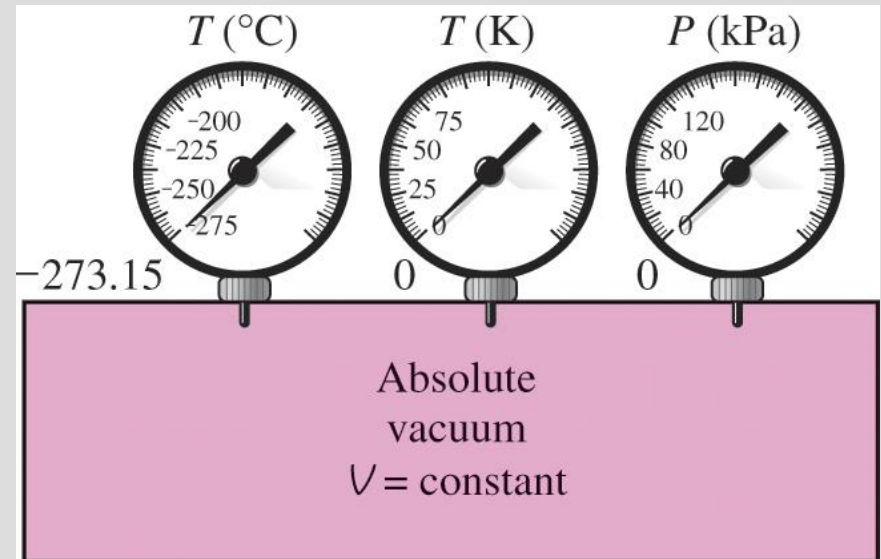
- All temperature scales are based on some easily reproducible states such as the freezing and boiling points of water: the *ice point* and the *steam point*.
- **Ice point:** A mixture of ice and water that is in equilibrium with air saturated with vapor at 1 atm pressure (0°C or 32°F).
- **Steam point:** A mixture of liquid water and water vapor (with no air) in equilibrium at 1 atm pressure (100°C or 212°F).
- **Celsius scale:** in SI unit system
- **Fahrenheit scale:** in English unit system

# Thermodynamic Temperature Scale

- **Thermodynamic temperature scale:** A temperature scale that is independent of the properties of any substance.
- **Kelvin scale** (SI) **Rankine scale** (E)
- A temperature scale nearly identical to the Kelvin scale is the **ideal-gas temperature scale**. The temperatures on this scale are measured using a **constant-volume gas thermometer**.



$P$  versus  $T$  plots of the experimental data obtained from a constant-volume gas thermometer using four different gases at different (but low) pressures.



A constant-volume gas thermometer would read  $273.15^{\circ}\text{C}$  at absolute zero pressure.



$$T(\text{K}) = T(^{\circ}\text{C}) + 273.15$$

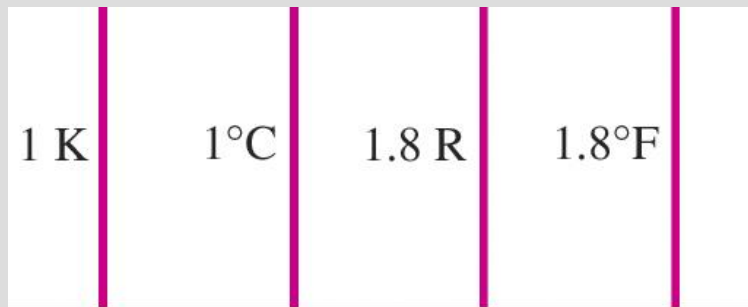
$$T(\text{R}) = T(^{\circ}\text{F}) + 459.67$$

$$T(\text{R}) = 1.8T(\text{K})$$

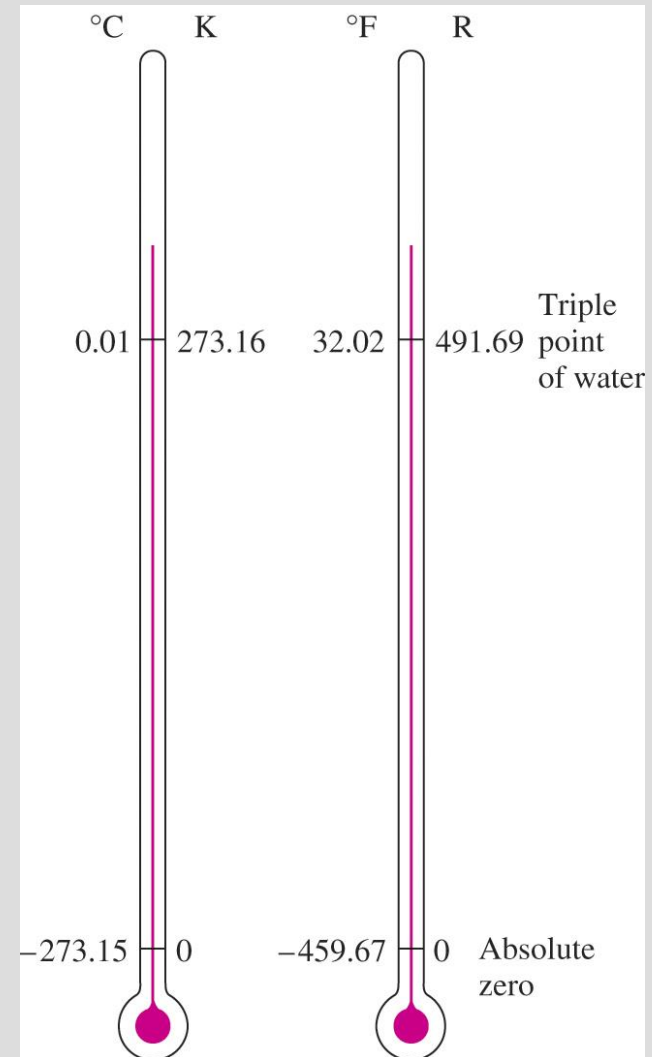
$$T(^{\circ}\text{F}) = 1.8T(^{\circ}\text{C}) + 32$$

$$\Delta T(\text{K}) = \Delta T(^{\circ}\text{C})$$

$$\Delta T(\text{R}) = \Delta T(^{\circ}\text{F})$$



Comparison of magnitudes of various temperature units.



Comparison of temperature scales.

# PRESSURE

**Pressure:** A normal force exerted by a fluid per unit area

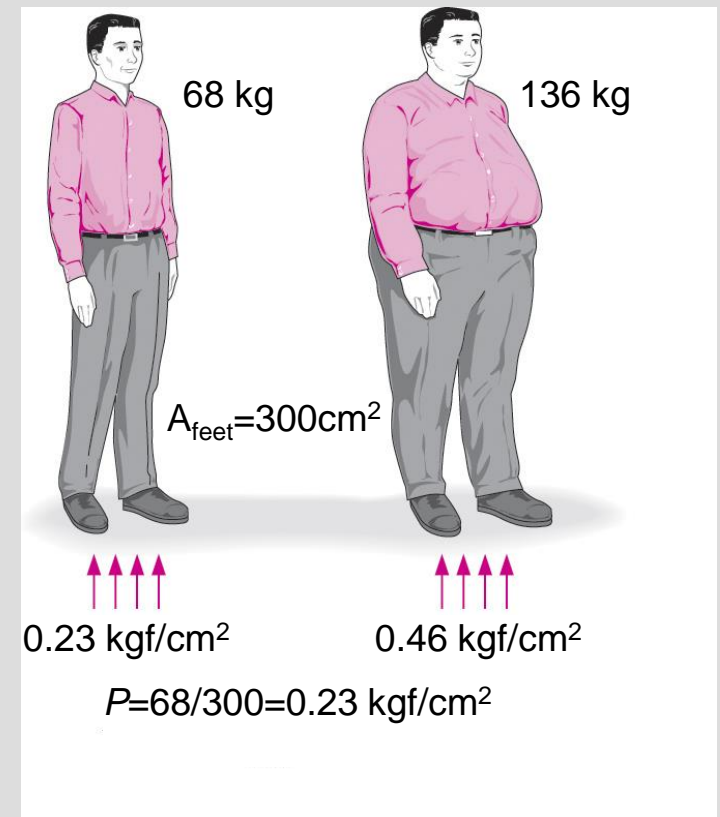
$$1 \text{ Pa} = 1 \text{ N/m}^2$$

$$1 \text{ bar} = 10^5 \text{ Pa} = 0.1 \text{ MPa} = 100 \text{ kPa}$$

$$1 \text{ atm} = 101,325 \text{ Pa} = 101.325 \text{ kPa} = 1.01325 \text{ bars}$$



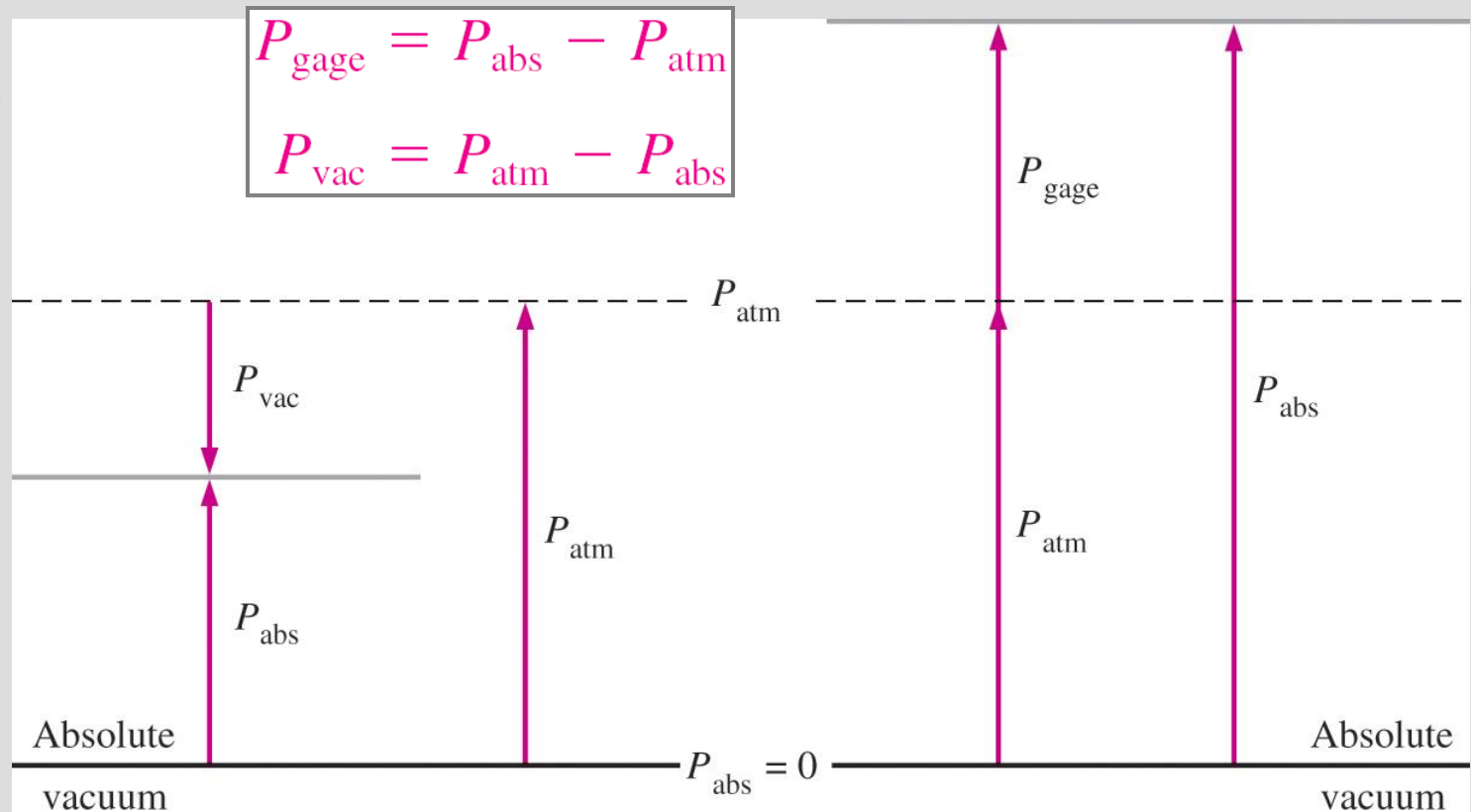
Some basic pressure gages.



The normal stress (or “pressure”) on the feet of a chubby person is much greater than on the feet of a slim person.

- **Absolute pressure:** The actual pressure at a given position. It is measured relative to absolute vacuum (i.e., absolute zero pressure).
- **Gage pressure:** The difference between the absolute pressure and the local atmospheric pressure. Most pressure-measuring devices are calibrated to read zero in the atmosphere, and so they indicate gage pressure.
- **Vacuum pressures:** Pressures below atmospheric pressure.

Throughout this text, the pressure  $P$  will denote **absolute pressure** unless specified otherwise.



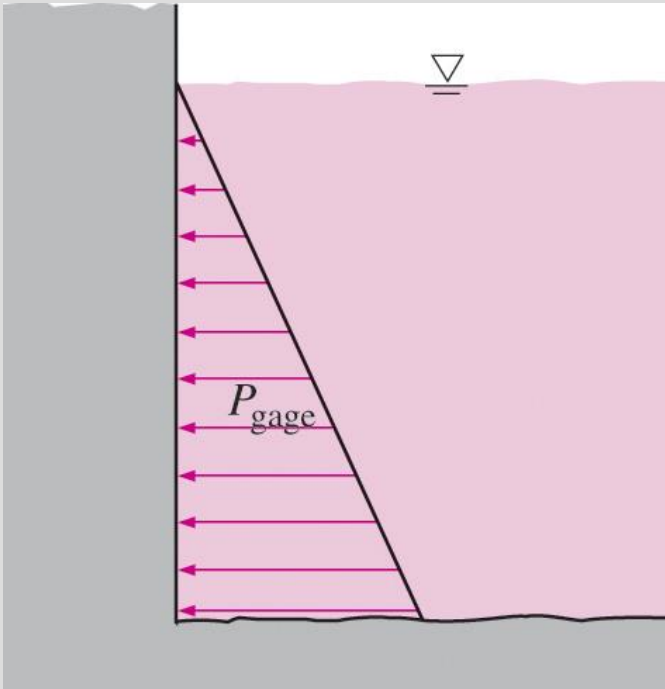
# Variation of Pressure with Depth

$$\Delta P = P_2 - P_1 = \rho g \Delta z = \gamma_s \Delta z$$

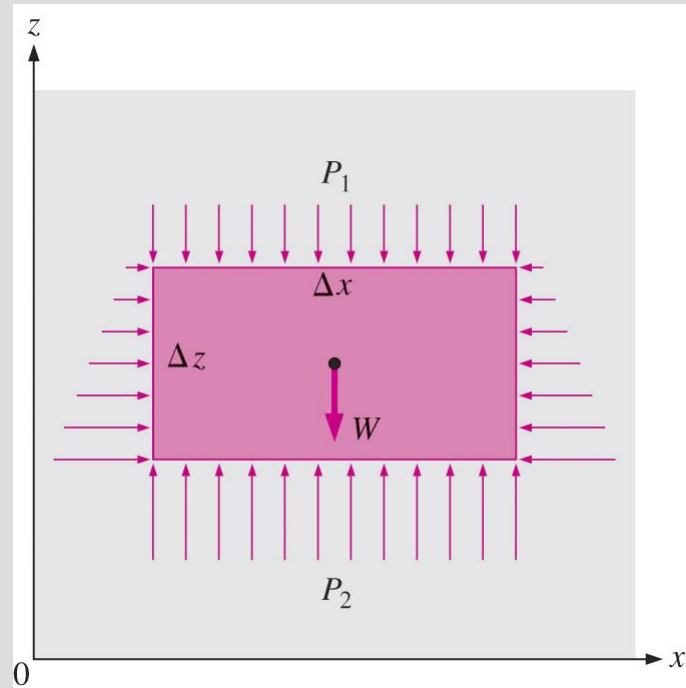
$$P = P_{\text{atm}} + \rho g h \quad \text{or} \quad P_{\text{gage}} = \rho g h$$

When the variation of density with elevation is known

$$\Delta P = P_2 - P_1 = - \int_1^2 \rho g dz$$

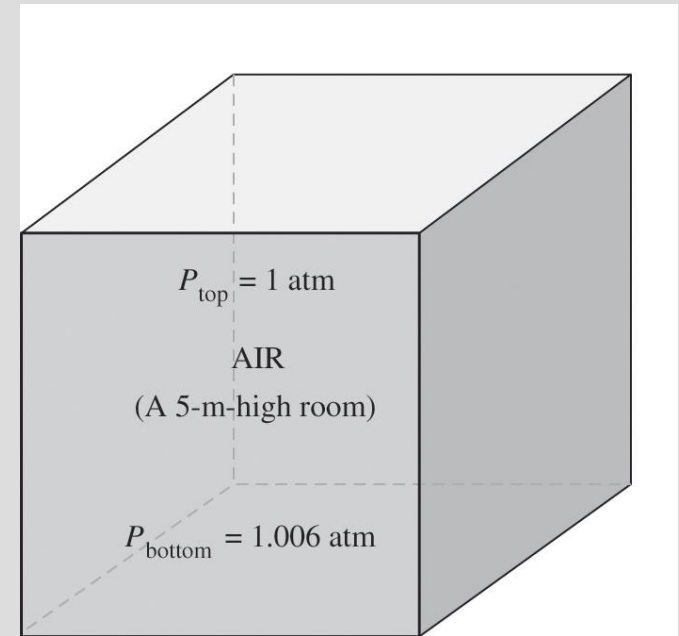


The pressure of a fluid at rest increases with depth (as a result of added weight).

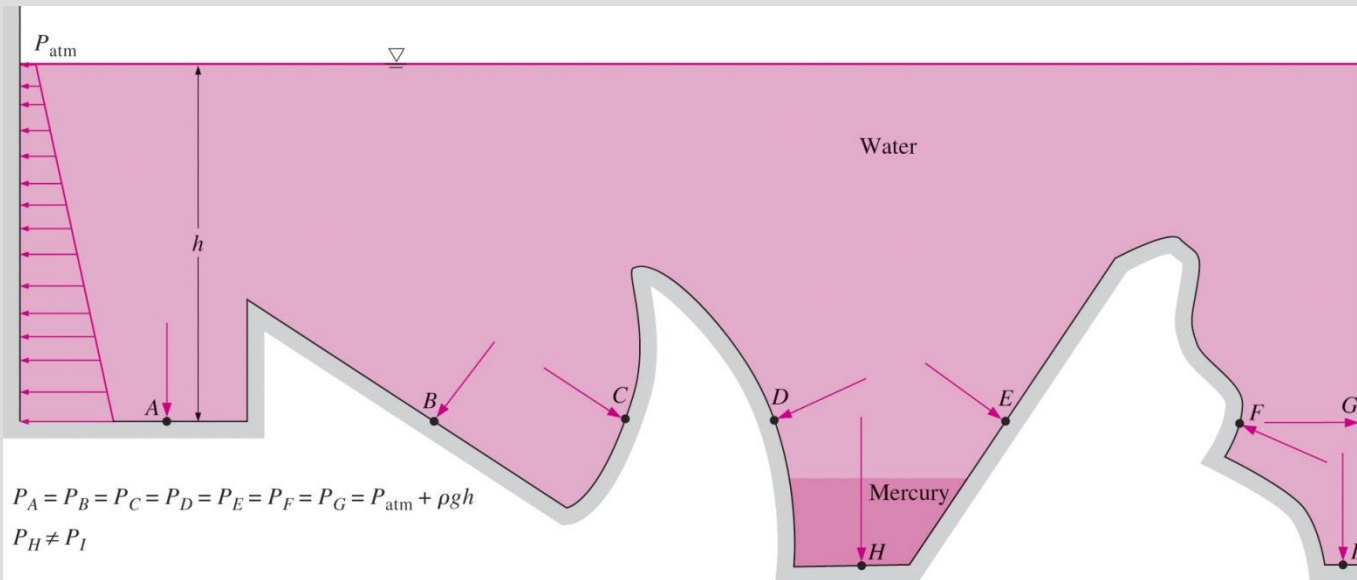
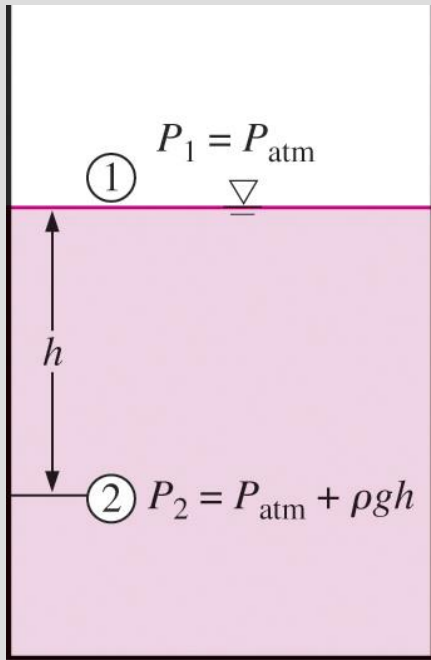


Free-body diagram of a rectangular fluid element in equilibrium.

In a room filled with a gas, the variation of pressure with height is negligible.



Pressure in a liquid at rest increases linearly with distance from the free surface.



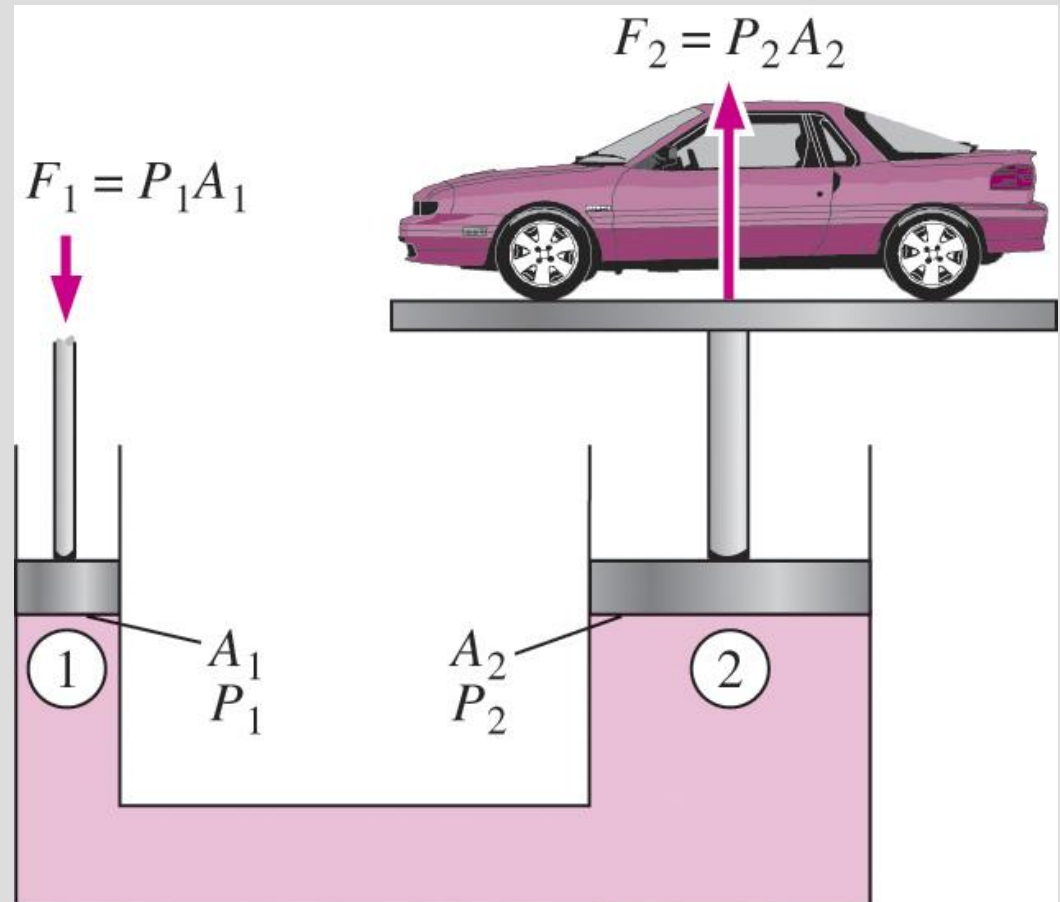
The pressure is the same at all points on a horizontal plane in a given fluid regardless of geometry, provided that the points are interconnected by the same fluid.

**Pascal's law:** The pressure applied to a confined fluid increases the pressure throughout by the same amount.

$$P_1 = P_2 \rightarrow \frac{F_1}{A_1} = \frac{F_2}{A_2} \rightarrow \frac{F_2}{F_1} = \frac{A_2}{A_1}$$

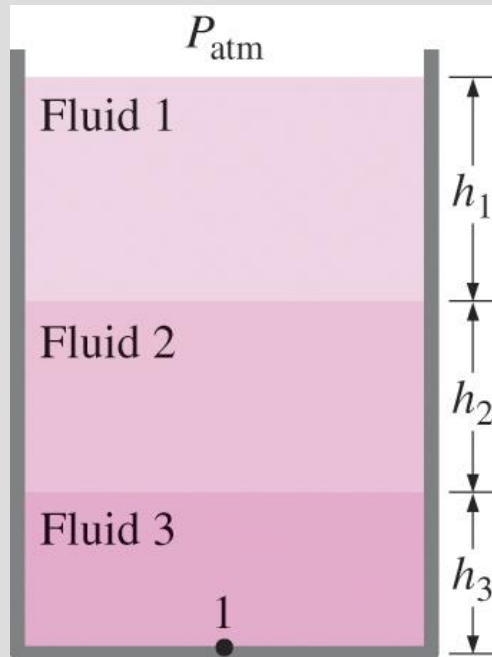
The area ratio  $A_2/A_1$  is called the *ideal mechanical advantage* of the hydraulic lift.

Lifting of a large weight by a small force by the application of Pascal's law.



# The Manometer

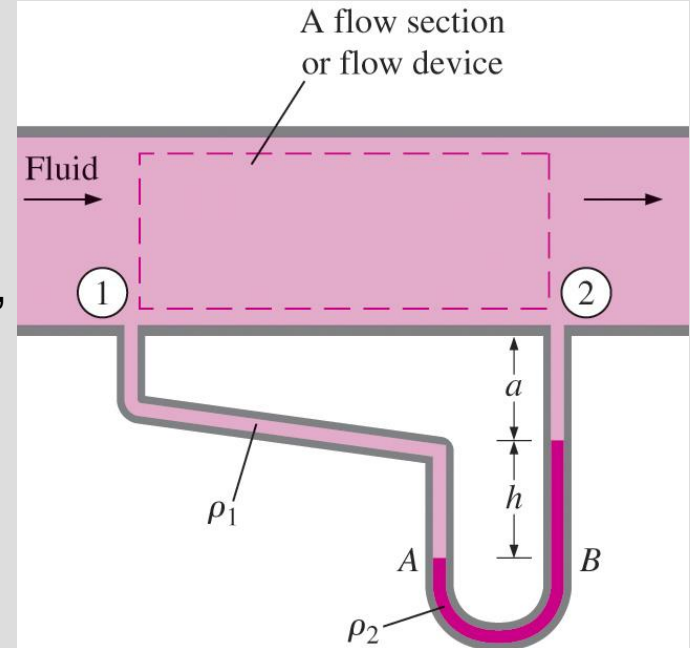
It is commonly used to measure small and moderate pressure differences. A manometer contains one or more fluids such as mercury, water, alcohol, or oil.



$$P_{\text{atm}} + \rho_1 g h_1 + \rho_2 g h_2 + \rho_3 g h_3 = P_1$$

In stacked-up fluid layers, the pressure change across a fluid layer of density  $\rho$  and height  $h$  is  $\rho g h$ .

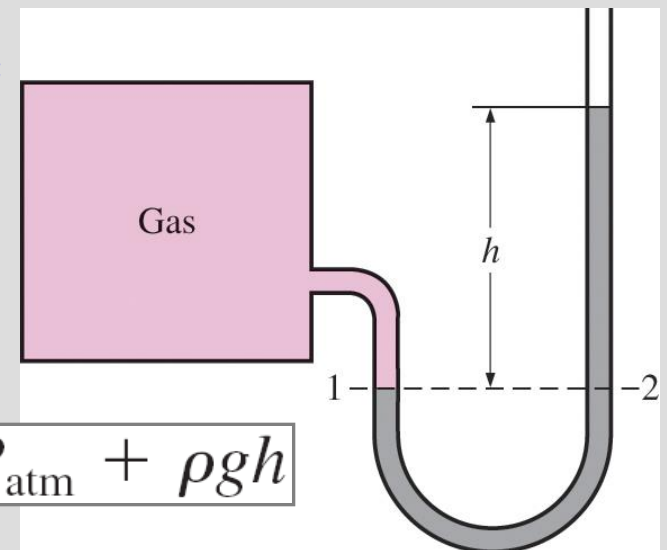
Measuring the pressure drop across a flow section or a flow device by a differential manometer.



$$P_1 + \rho_1 g (a + h) - \rho_2 g h - \rho_1 g a = P_2$$

$$P_1 - P_2 = (\rho_2 - \rho_1) g h$$

The basic manometer.

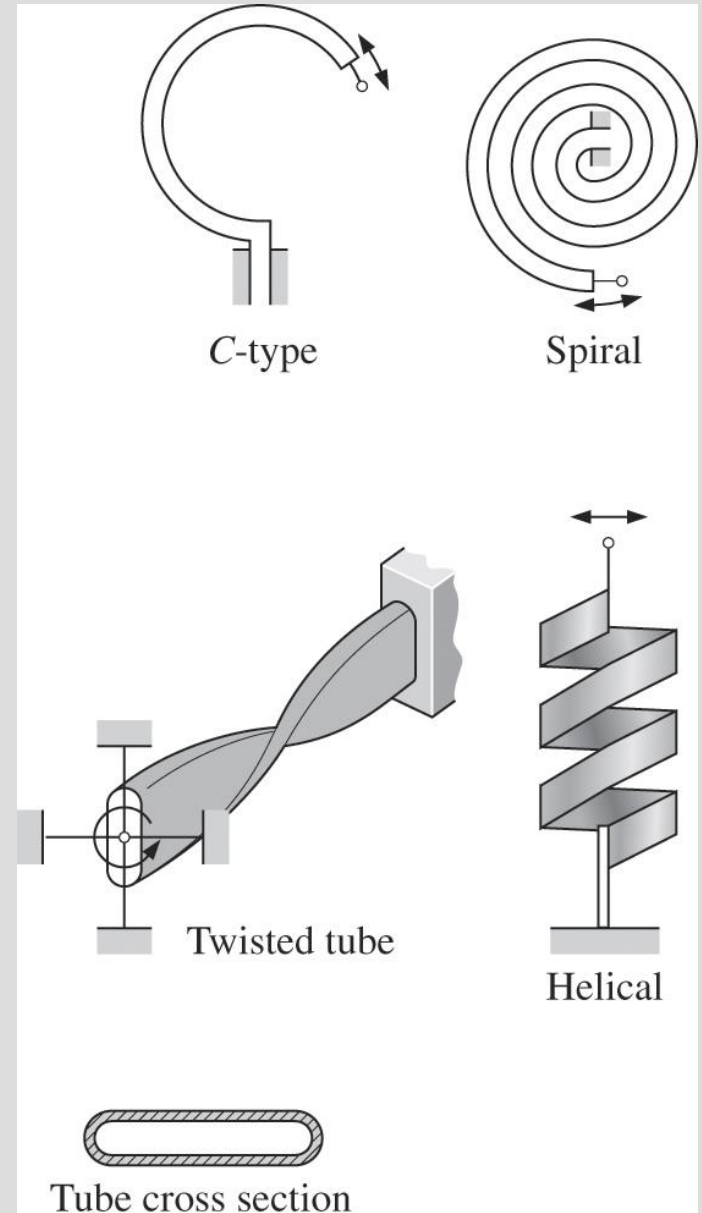


$$P_2 = P_{\text{atm}} + \rho g h$$

# Other Pressure Measurement Devices

- **Bourdon tube:** Consists of a hollow metal tube bent like a hook whose end is closed and connected to a dial indicator needle.
- **Pressure transducers:** Use various techniques to convert the pressure effect to an electrical effect such as a change in voltage, resistance, or capacitance.
- Pressure transducers are smaller and faster, and they can be more sensitive, reliable, and precise than their mechanical counterparts.
- **Strain-gage pressure transducers:** Work by having a diaphragm deflect between two chambers open to the pressure inputs.
- **Piezoelectric transducers:** Also called **solid-state pressure transducers**, work on the principle that an electric potential is generated in a crystalline substance when it is subjected to mechanical pressure.

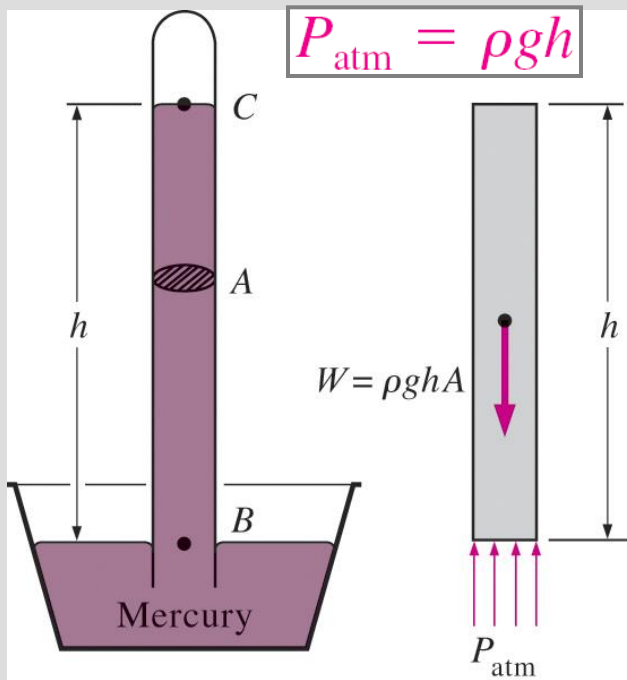
Various types of Bourdon tubes used to measure pressure.



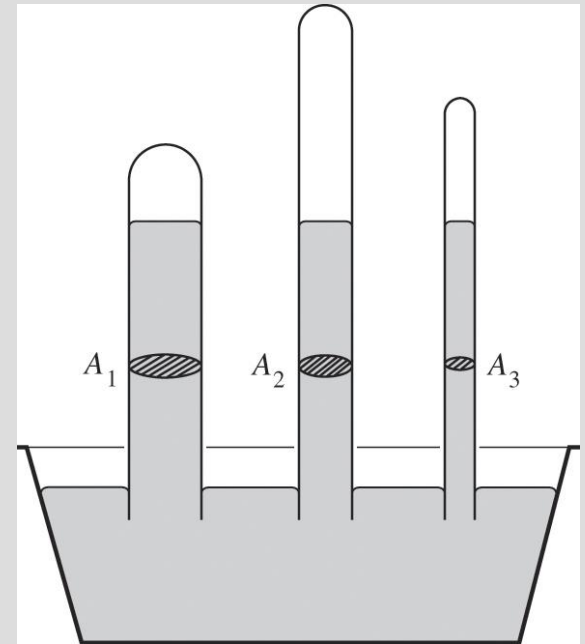


# THE BAROMETER AND ATMOSPHERIC PRESSURE

- Atmospheric pressure is measured by a device called a **barometer**; thus, the atmospheric pressure is often referred to as the **barometric pressure**.
- A frequently used pressure unit is the **standard atmosphere**, which is defined as the pressure produced by a column of mercury 760 mm in height at 0°C ( $\rho_{\text{Hg}} = 13,595 \text{ kg/m}^3$ ) under standard gravitational acceleration ( $g = 9.807 \text{ m/s}^2$ ).



The length or the cross-sectional area of the tube has no effect on the height of the fluid column of a barometer, provided that the tube diameter is large enough to avoid surface tension (capillary) effects.



The basic barometer.

# PROBLEM-SOLVING TECHNIQUE

- Step 1: Problem Statement
- Step 2: Schematic
- Step 3: Assumptions and Approximations
- Step 4: Physical Laws
- Step 5: Properties
- Step 6: Calculations
- Step 7: Reasoning, Verification, and Discussion

## **EES (Engineering Equation Solver)** (Pronounced as ease):

EES is a program that solves systems of linear or nonlinear algebraic or differential equations numerically. It has a large library of built-in thermodynamic property functions as well as mathematical functions. Unlike some software packages, EES does not solve engineering problems; it only solves the equations supplied by the user.

# Summary

- Thermodynamics and energy
  - ✓ Application areas of thermodynamics
- Importance of dimensions and units
  - ✓ Some SI and English units, Dimensional homogeneity, Unity conversion ratios
- Systems and control volumes
- Properties of a system
- Density and specific gravity
- State and equilibrium
  - ✓ The state postulate
- Processes and cycles
  - ✓ The steady-flow process
- Temperature and the zeroth law of thermodynamics
  - ✓ Temperature scales
- Pressure
  - ✓ Variation of pressure with depth
- The manometer and the atmospheric pressure
- Problem solving technique