**Lab sheet #2**

**Preparation of solutions:**

**Method:**

(1) **You are provided with solid NaOH, Prepare 50ml with 0.08M NaOH solution.**

Calculation:

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🡺To prepare the 0.08M NaOH solution, …………….g of solid NaOH should be dissolved in a little volume of water then the volume made up to ………….ml ,by the addition of water.

(2) **You are provided with solid NaCl, Prepare 50ml with 1.5 w/v% solution of NaCl.**

Calculation:

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🡺 To prepare the 1.5 w/v% solution, …………. g of NaCl should be dissolved in little water and the volume made up to ……….ml by the addition of water.

(3) **Prepare 100ml with 0.4 M HCl solutions starting with the concentrated HCl solution you are provided with: (w/w%= 36, S.Gr =1.18 ).**

Calculation:

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🡺 To prepare the 100ml of 0.4M HCl solution ……….ml of stock (i.e. concentrated HCl) solution is needed and the volume made up to ……..ml by the addition of water.

**Note:** Atomic weights: Na = 23, Cl= 35.5, O = 16, H = 1