**BCH 312**

**Lab sheet #6**

**Name:**……………………………………………………………………………… **ID:** ………………………………………………………………………………

**Method and Calculations**:

You are provided with 10 ml of a 0.1M CH3COOH weak acid solution, titrate it with 0.1M NaOH adding the base drop wise mixing, and recording the pH after each 0.5 ml NaOH added until you reach a pH=10.[ stop until you add 10 ml NaOH]

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| **Measured pH value** | **Amount of 0.1M NaOH added [ml]** |
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[1] record the values in titration table and Plot a Curve of pH versus ml of NaOH added.

[2]Calculate the pH of the weak acid HA solution after the addition of 3ml, 5ml, and 10ml of NaOH.

[3] determine the Pka value of weak acid.

[4]Compare your calculated pH values with those obtained from Curve.