**Lab sheet #8**

**Method:**

* You are provided with:

1. Standard solution (Stock solution) of **Copper Sulfate with 0.1 M** [known concentration].
2. Solution with **Unknown concentration "A".**
3. Solution with **Unknown concentration "B".**

* Set up 8 test tubes, as following table:

|  |  |  |  |
| --- | --- | --- | --- |
| **Tube** | **0.1M Copper Sulfate Standard Solution (ml)** | **Distal water (ml)** | **Solutions with unknown concentration (ml)** |
| **Blank** | **-** | 10 ml | - |
| **A** | 2ml | 8ml | - |
| **B** | 4ml | 6ml | - |
| **C** | 6ml | 4ml | - |
| **D** | 8ml | 2ml | - |
| **E** | 10ml | - | - |
| **Solution "A"** | - | - | 3 ml |
| **Solution "B"** | - | - | 3 ml |

* Mix the contents using the vortex.
* Measure the absorbance of each tube at 600 nm against the blank [………..….].

**Results:**

|  |  |  |
| --- | --- | --- |
| **Tube** | **Absorbance at 600nm** | **Concentration (M)** C1 x V1 = C2 x V2 |
| **A** |  |  |
| **B** |  |  |
| **C** |  |  |
| **D** |  |  |
| **E** |  |  |
| **Solution "A"** |  | From the curve= |
| **Solution "B"** |  | From the curve= |

* Calculate the concentrations of the series of known standard solutions.
* Plot the standard curve (Absorbance vs. Concentration of standard solutions).
* Determine the concentration of Solution "A" and "B" using the graph and Beer’s- Lambert law.