

	n-Butane	C ₄ H ₁₀
Boiling Point =	30.9	° F
R =	8.3145	J mole ⁻¹ K ⁻¹
L _v =	23186.65	J/ mole
L _v =	23.19	KJ/ mole

Slope =	-2788.7	Ln(14.65) =	2.68
Intercept =	12.923	1/T =	0.0037 °K
		T =	272.4 °K
		T =	-0.6 °C
		T =	30.9 °F

T, °F	P°, psia	T, °C	T, °K	1/T (°K)	Ln(P)
0	7.3	-17.8	255.2	0.0039	1.99
10	9.2	-12.2	260.8	0.0038	2.22
20	11.6	-6.7	266.3	0.0038	2.45
30	14.4	-1.1	271.9	0.0037	2.67
40	17.7	4.4	277.4	0.0036	2.87
50	21.6	10.0	283.0	0.0035	3.07
60	26.3	15.6	288.6	0.0035	3.27
70	31.6	21.1	294.1	0.0034	3.45
80	37.6	26.7	299.7	0.0033	3.63
90	44.5	32.2	305.2	0.0033	3.80
100	52.2	37.8	310.8	0.0032	3.96
110	60.8	43.3	316.3	0.0032	4.11
120	70.8	48.9	321.9	0.0031	4.26
130	81.4	54.4	327.4	0.0031	4.40
140	92.6	60.0	333.0	0.0030	4.53



