



Boiling Point = 96.0

°F

R = 8.3145

$\text{J mole}^{-1} \text{K}^{-1}$

L = 25529.67

J/mole

L = 25.53

KJ/mole

Formula

Slope = -3070.5

Intercept = 12.636

$Y = -3070.5 X + 12.636$

$\ln(14.65) = 2.68$

$1/T = 0.0032$ °K

T = 308.5 °K

T = 35.5 °C

T = 96.0 °F

T, °F	P°, psia	T, °R	1/T (°R)	T, °C	T, °K	1/T (°K)	Ln(P)
60	7.0	520	0.00192	15.6	288.6	0.0035	1.95
100	15.6	560	0.00179	37.8	310.8	0.0032	2.75
110	19.0	570	0.00175	43.3	316.3	0.0032	2.94
120	22.5	580	0.00172	48.9	321.9	0.0031	3.11
130	26.6	590	0.00170	54.4	327.4	0.0031	3.28
140	31.0	600	0.00167	60.0	333.0	0.0030	3.43
150	36.6	610	0.00164	65.6	338.6	0.0030	3.60
160	42.3	620	0.00161	71.1	344.1	0.0029	3.74
170	48.9	630	0.00159	76.7	349.7	0.0029	3.89
180	54.1	640	0.00156	82.2	355.2	0.0028	3.99
190	57.4	650	0.00154	87.8	360.8	0.0028	4.05
200	67.5	660	0.00152	93.3	366.3	0.0027	4.21

